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Chemistry Chapter 5 Review

Chemistry: Chapter 5 Review. STUDY. Flashcards. Learn. Write. Spell. Test. PLAY. Match. Gravity. Created by. cherren. Chemistry: Matter and Change. Terms in this set (62) Amplitude. The height of a wave from the origin to a crest, or from the origin to a trough. Atomic Emission Spectrum.

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Chemistry 1: Chapter 5 Review. Magnesium (Mg) Phosphorus (P) Lead (Pb) Lawrencium (Lr) Determine the element: 1s2 2s2 2p6 3s2. Determine the element: 1s2 2s2 2p6 3s2 3p3. What is this element: 1s2 2s2 2p6 3s2 3p6 4s2 3d10 4p6 5s2 4d1…. Determine the element: [Rn] 7s2 5f14.

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Chemistry: Chapter 5 Review. Chemistry: Chapter 5 Review. Describe…. 1) Dalton’s model – solid sphere, made up of only 1 type of matter, no space. 2) Schrodinger’s model – based on probability of finding an e-; electrons are found in orbitals (look like electron clouds), which are in sublevels (s,p,d,f,g), which are in energy levels.

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Chapter 5 Chemistry review. electromagnetic radiation. Wavelength. frequency. Amplitude. a kind of radiation including visible light, radio waves, gamm…. Horizontal distance between the crests or between the troughs…. the number of complete wavelengths that pass a point in a give…. Height of a wave.

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Chemistry Chapter 5 Review 1. What is region of high probability of finding an electron called? 2. What is the lowest energy level? 3. What is the tendency of electrons to enter orbitals of lowest energy first called? 4. The arrangement of electrons around atomic nucleus is what? 5. What orbital has at most two electrons? 6.

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CHAPTER 5 REVIEW The Periodic Law SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. c When an electron is added to a neutral atom, energy is (a) always absorbed. (c) either absorbed or released. (b) always released. (d) neither absorbed nor released. 2. d The energy required to remove an electron from a neutral atom is the atom’s

5 The Periodic Law
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CHAPTER 4 REVIEW Arrangement of Electrons in Atoms MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Under what conditions is a photon emitted from an atom? A photon is emitted when an electron moves from a higher energy level to a lower energy level. 2. What do quantum numbers describe?

4 Arrangement of Electrons in Atoms
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- Chapter 5 & 6 Review 1.6.20 The scientists you need to know and what they did Rutherford proposed that all of an atom’s positive charge and virtually all of its mass are in a nucleus surrounded by fast-moving electrons(his nuclear model). Boseman Suggested that molecules could get a discrete packet of energy instead of continuous light Bohr Made the current model for atoms Said that ...